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9-1 Introduction to Stoichiometry pages

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275-277 Questions # 1-3. 9-2 Ideal
Stoichiometric Calculations pages

280-287 Questions # 1ab,2a,3a . 9-3
Limiting Reactants and Percent Yield

pages 288-294 Questions # 1-2 EOC's
Page 295

#2,7,10a,12ab,17a,22a,28a,33.

Objectives: By the end of this unit you
should... Define Stoichiometry.

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Chapter 9 Stoichiometry - PC\|MAC

stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry is the

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subject of this chapter and it is based on

CHAPTER 9 Stoichiometry - Quia

CHAPTER 9 DO NOT EDIT--Changes must be made through "File info" ... Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction stoichiometry calculations start

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with a balanced chemical equation. This equation gives the

**CorrectionKey=NL-A DO NOT
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- 9.1. Objectives.
- Define stoichiometry.
 - Describe the importance of the mole ratio in stoichiometric calculations.
 - Write a mole ratio relating two

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substances in a chemical equation.

2/26/2018 2. Stoich Definitions.

- Composition stoichiometry deals with the mass relationships of elements in compounds.
- Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.

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Chapter 9

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reaction. Reaction stoichiometry is the

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Chapter 9 Review Stoichiometry Worksheet Answers

Chapter 9: Standard Review Worksheet

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1. Answers will vary. An example is included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This describes the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

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Chapter 9: Standard Review Worksheet

CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the

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percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

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Chapter 9 - Stoichiometry - Ms. Clark's Website

Stoichiometry Chapter 9, p. 275 - 294
Intro to Stoichiometry • Reaction

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Stoichiometry: Involves the mass relationships between reactants and products in a chemical reaction •

Coefficients in a chemical reaction represent the mole ratios of each substance that react together •

Example: $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$

Chapter 9 Stoichiometry Notes -

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9.3 Objectives Describe a method for determining which of two reactants is a limiting reactant. Calculate the amount in moles or mass in grams of a product, given the amounts in moles or masses in grams of two reactants, one of which is in excess.; Distinguish between theoretical yield, actual yield, and

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percentage yield.; Calculate percentage yield, given the actual yield and

Chapter 9: Stoichiometry - HHS Chemistry

Chapter 9 - Stoichiometry - Ms. Clark's Website Stoichiometry is the branch of chemistry that deals with elements in compounds and with reactants and

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products in chemical reactions, focusing on... the law of conservation of mass
Reaction Stoichiometry is based on chemical equations and... the number of significant figures of any measured quantities in the problem
Stoichiometry
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Chapter 9 Stoichiometry Section 1 Answers

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Chapter 03 - Stoichiometry

5.0 g Cu 1 mol Cu 1 mol Ag 107.9 g Ag =
8.5 g Ag. 63.5 g Cu 1 mol Cu 1 mol Cu.

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$8.5 \times 100 = 55.9\%$ yield. 15.2 CHAPTER 11: STOICHIOMETRY. MOLE TO MOLE RATIO. When nitrogen and hydrogen gas are heated under the correct conditions, ammonia gas (NH_3) is formed. a. RXN: $1. \text{N}_2 + 3. \text{H}_2 \rightarrow 2. \text{NH}_3$. b.

CHAPTER 11: STOICHIOMETRY

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