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specifies that we can get molecular orbitals through a linear combination of atomic orbitals; that is, by adding and subtracting them.

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its center at the nucleus of the atom. P orbital has dumb-bell shape and it has 3 orbitals p_x , p_y , p_z . Each one contains two lobes and nucleus lies between them.

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often discussed in terms of localized molecular orbitals in a qualitative and informal sense.

Historically, much of classical organic chemistry was built on the older valence bond / orbital hybridization models of bonding. To account for phenomena like aromaticity, this simple model of bonding is supplemented by semi-quantitative results

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Organic chemists with a serious interest in understanding and explaining their work usually express their ideas in molecular orbital terms, so much

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so that it is now an essential component of every organic chemist's skills to have some acquaintance with molecular orbital theory. Molecular Orbitals and Organic Chemical Reactions is both a simplified account of molecular orbital theory and a review of its applications in organic chemistry; it provides a basic introduction to the subject and a ...

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Molecular Orbitals

When two hydrogen atoms come together to form the hydrogen molecule, the atomic s orbitals of each atom are combined to form two molecular orbitals.

One of these new orbitals is the result of the addition of the two atomic orbitals, while

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the other is created by
a subtraction of these
orbitals.

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Draw a molecular orbital energy diagram when the 1s orbitals of two hydrogen atoms combine to form a hydrogen molecule. In your drawing, label the molecular orbitals formed and show properly how the electrons are

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distributed in the
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molecule with 6 carbon
atoms each having a p
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Molecular orbital theory is used by chemists to describe the arrangement of electrons in chemical structures. It is also a theory capable of giving some insight

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into the forces involved in the making and breaking of chemical bonds—the chemical reactions that are often the focus of an organic chemist's interest.

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